$C_2H_5OH_{(L)} + 3O_{2(G)} \rightarrow 2CO_{2(G)} + 3H_2O_{(G)}$

 When 73.0 moles of ethanol combusts, how many moles of oxygen are consumed?

 $C_2H_5OH_{(L)} + 3O_{2(G)} \rightarrow 2CO_{2(G)} + 3H_2O_{(G)}$

2. If 1.75 x 10^{24} molecules of ethanol combusts, how many liters of CO_{2(G)} form?

ZnCl_{2(AQ)} + 2K_(S) → 2KCl_(AQ) + Zn_(S) 3. When 33.3 moles of potassium react, how many grams of zinc will be produced?

 $ZnCl_{2(AQ)} + 2K_{(S)} \rightarrow 2KCl_{(AQ)} + Zn_{(S)}$

4. When 12.43 moles of zinc chloride react, how many moles of potassium chloride form?

$N_{2(G)} + 3H_{2(G)} \rightarrow 2NH_{3(G)}$

5. When 358 liters of hydrogen react, how many grams of ammonia will form?

 $N_{2(G)} + 3H_{2(G)} \rightarrow 2NH_{3(G)}$

6. When 0.456 moles of nitrogen react, how many molecules of ammonia will form?