

Balance each of these reactions, with phase symbols.

Name: _____

Skeleton	Type of reaction
$\text{H}_{2(\text{G})} + \text{O}_{2(\text{G})} \rightarrow \text{H}_2\text{O}_{(\text{G})}$	
$\text{Sr}(\text{OH})_{2(\text{AQ})} + \text{Li}_2\text{CrO}_{4(\text{AQ})} \rightarrow$	
$\text{ZnBr}_{2(\text{AQ})} + \text{Al}_{(\text{S})} \rightarrow$	
$\text{C}_{(\text{S})} + \text{S}_{8(\text{S})} \rightarrow \text{CS}_{2(\text{S})}$	
$\text{Na}_{(\text{S})} + \text{O}_{2(\text{G})} \rightarrow \text{Na}_2\text{O}_{2(\text{S})}$	
$\text{N}_{2(\text{G})} + \text{O}_{2(\text{G})} \rightarrow \text{N}_2\text{O}_{5(\text{G})}$	
$\text{P}_{(\text{S})} + \text{Cl}_{2(\text{G})} \rightarrow \text{PCl}_{5(\text{S})}$	
$\text{Na}_{(\text{S})} + \text{O}_{2(\text{G})} \rightarrow \text{Na}_2\text{O}_{(\text{S})}$	
$\text{Al}_{(\text{S})} + \text{S}_{8(\text{S})} \rightarrow \text{Al}_2\text{S}_{3(\text{S})}$	
$\text{H}_2\text{O}_{(\text{L})} \rightarrow \text{H}_{2(\text{G})} + \text{O}_{2(\text{G})}$	
$\text{Mg}_{(\text{S})} + \text{Cl}_{2(\text{G})} \rightarrow \text{MgCl}_{2(\text{S})}$	
$\text{C}_{15}\text{H}_{32(\text{S})} + \text{O}_{2(\text{G})} \rightarrow \text{CO}_{2(\text{G})} + \text{H}_2\text{O}_{(\text{G})}$	
$\text{C}_6\text{H}_{6(\text{G})} + \text{O}_{2(\text{G})} \rightarrow \text{CO}_{2(\text{G})} + \text{H}_2\text{O}_{(\text{G})}$	
$\text{N}_{2(\text{G})} + \text{H}_{2(\text{G})} \rightarrow \text{NH}_{3(\text{G})}$	

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Skeleton	Type of reaction
$\text{Li}_{(s)} + \text{AlCl}_{3(AQ)} \rightarrow$	
$\text{C}_2\text{H}_6(G) + \text{O}_2(G) \rightarrow \text{CO}_2(G) + \text{H}_2\text{O}(G)$	
$\text{Rb}_{(s)} + \text{P}_{(s)} \rightarrow \text{Rb}_3\text{P}_{(s)}$	
$\text{CH}_4(G) + \text{O}_2(G) \rightarrow \text{CO}_2(G) + \text{H}_2\text{O}(G)$	
$\text{Na}_{(s)} + \text{I}_{2(s)} \rightarrow \text{NaI}_{(s)}$	
$\text{Rb}_{(s)} + \text{S}_{8(s)} \rightarrow \text{Rb}_2\text{S}_{(s)}$	
$\text{NH}_3(AQ) + \text{HCl}(AQ) \rightarrow \text{NH}_4\text{Cl}(AQ)$	
$\text{Li}_{(s)} + \text{SnCl}_4(AQ) \rightarrow$	
$\text{NH}_3(G) \rightarrow \text{N}_2(G) + \text{H}_2(G)$	
$\text{Cs}_{(s)} + \text{N}_2(G) \rightarrow \text{Cs}_3\text{N}_{(s)}$	
$\text{CaCO}_3(s) \rightarrow \text{CaO}_{(s)} + \text{CO}_2(G)$	
$\text{C}_{10}\text{H}_{22(s)} + \text{O}_2(G) \rightarrow \text{CO}_2(G) + \text{H}_2\text{O}(G)$	
$\text{C}_{(s)} + \text{O}_2(G) \rightarrow \text{CO}_2(G)$	
$\text{C}_3\text{H}_8(G) + \text{O}_2(G) \rightarrow \text{CO}_2(G) + \text{H}_2\text{O}(G)$	