

During this lab we will examine several relationships between pressure, volume, temperature - and mass of gases.



In 1662, Robert Boyle wrote:

“For a fixed amount of an ideal gas kept at fixed temperature, Pressure and Volume are inversely proportional.”
(At constant temperature, pressure + volume are inversely proportional.)

In the 1780's Jacques Charles of France wrote that

“At constant pressure, the volume of a given mass of an ideal gas increases or decreases by the same factor at its temperature on the absolute scale.”

(At constant pressure, Volume and Temperature is directly proportional.)

In 1802, the French chemist Joseph Louis Gay-Lussac

“The pressure for a fixed mass of gas is directly proportional to the gas's temperature.”
(At constant volume, pressure + temperature are directly proportional.)

The author of this lab is Dr. Arden P. Zipp, of SUNY Cortland. He was my teacher. He is the man that started my love and understanding of chemistry. He would say he doesn't belong on this page, but I think he does, sort of.

You can do these three experiment in any order. For safety, do not pump up the soda bottles with the bicycle pumps past 65 psi (pounds per square inch) on the pump gauge.

ALWAYS Point the bottles away from your classmates when pumping them up, and keep your goggles on at all times! High Pressure is invisible but can be dangerous, especially to eyes.

Write the combined gas law big in this space.

Review Guide	formula	Show a simple graph of this relationship	Name this relationship as <i>inversely or directly proportional</i> . Or both.
The combined gas law:			
The combined gas law with constant temp:			
The combined gas law with constant pressure:			
The combined gas law with constant volume:			

Experiment 1 Pressure + Mass of gases

Next, pump up the bottle to about 65 psi, check actual pressure with the hand held gauge, then mass the bottle.

Trial	Mass of bottle + gas (in grams)	Pressure in psi
1		
2		
3		
4		
5		
6		

Experiment 2 - Pressure and Volume are Inversely Proportional

Obtain a bottle with special cap with valve stem inserted. Get a 10 cm³ syringe with the rubber cap used to seal the syringe (it's labeled 10 cc). Get a bicycle pump. Carefully measure syringe open to exactly 10 cm³, and put on the rubber cap. Put syringe into soda bottle and close up. Pump up the soda bottle to a maximum of 65 psi on the bicycle gauge. Remove pump from bottle.

Slowly remove some air from the bottle, ONLY enough air is removed that you see the syringe move. Measure the pressure in the bottle again, and the corresponding syringe volume. Repeat ten times. If you run out of air pressure before ten trials, re-pump the bottle to get enough readings.

Make sure that the syringe volume changes a little bit each time you measure, and remember to reset the hand pressure gauge each time by pressing the button. Put data below.

Trial	Measured Bottle Pressure psi	Syringe Volume cm ³	P x V =
1			
2			
3			
4			
5			
6			
7			
8			
9			
10			

AVERAGE CONSTANT with units! (2 SF)



Experiment 3 Pressure + Temperature are directly proportional

Pump up the bottle to 65 psi on the bicycle pump as fast as you can. Your partner shall hold the bottle like a football, pointing the cap away from human beings. Note any temperature changes.

Once at the MAXIMUM of 65 psi detach bottle from pump. While pointing away from people, let the air out as fast as you can by twisting the cap ONE HALF TURN, as shown by teacher. Watch for the cloud!

Increasing pressure makes the bottle temperature HOTTER or COLDER ?

Decreasing pressure makes the bottle temperature HOTTER or COLDER ?

Lab Questions

1. Draw a small graph showing pressure as a function of volume, the title is:
The relationship between pressure + volume
2. Draw a small graph showing pressure as a function of temperature, the title is:
The relationship between pressure + temperature
3. Draw a small graph showing volume as a function of temperature, the title is:
The relationship between volume + temperature
4. Draw a small graph showing mass as a function of pressure, the title is:
The relationship between mass + pressure.
5. In experiment 3, did your P X V remain exactly the same for all ten trials? Should it have? Isn't $P \times V = a \text{ constant}$? Explain why your final column was not always the same.
6. Write out Avogadro's Hypothesis. In private, read it aloud 4 times, *with style*. If possible, orate it once at an appropriate volume in the commons during your lunch time (preferably while standing on a chair), or share it with another adult at school and help them to understand it. Really.
7. A 75.3 Liter sample of air at standard temperature and pressure is heated until the gas warms to 365 Kelvin. If the volume stays constant, what is the new gas pressure?
8. When a 75.3 Liter sample of air at standard temperature and pressure is heated, the gas warms to 365 Kelvin. If the volume expands to 91.2 Liters, what is the new gas pressure?

Make the 3 graphs about this size, use LABELS + TITLES

Y as a function of X

this lab	includes	points
cover page	Title, Define ideal gas and real gas.	1
8 questions	Do on loose leaf, 8 x 2 points each	16
On graph paper	Bottle Pressure as a function of syringe volume. (title, labels, units, best fit <u>curved</u> line)	2
last page	Explain what you measured and what you did in the 4 parts of the lab. List the 7 points of the KMT, explain each of these statements as the "real truth" or explain how it is "almost true". Tell under what conditions a real gas will act most ideally. Give an example of 3 different gases at the same conditions how one is more ideal than the other. Tell how when you compare 3 samples of oxygen gas, how one is more ideal than the other two samples.	6
	Due: _____	25 total possible