

REACTIONS HW #3 - Double Replacement Reactions name: _____

If you mix two aqueous solutions and a precipitate is formed, that is an indicator that a type of a double replacement reaction has occurred. BALANCE THESE EQUATIONS with phase symbols. Use table F to define AQ or S for each product. One "reaction" is not(?)

Sodium carbonate + zinc chlorate solutions react



Copper (II) sulfate + calcium chloride solutions react



Potassium hydroxide + lead (IV) nitrate solutions react



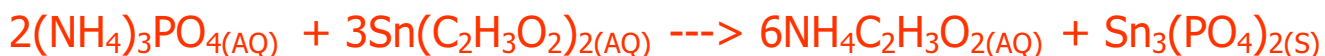
Silver hydrogen carbonate + Iron (II) bromide solutions react



Barium hydroxide + lithium sulfate solutions react



Ammonium phosphate + tin (II) acetate solutions react



Calcium chromate + sodium sulfide solutions react



Strontium acetate + lithium carbonate solutions react



Ammonium phosphate + rubidium nitrate solutions react



NO REACTION: When both products are AQ, it's really just a mixture!