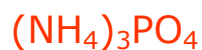


Mole HW #2 **ANSWERS**

Write the formula of each of these four compounds correctly, then find the MOLR MASS.

Ammonium Phosphate



$$N \quad 3 \times 14 = 42$$

$$H \quad 12 \times 1 = 12$$

$$P \quad 1 \times 31 = 31$$

$$O \quad 4 \times 16 = \underline{64}$$

149 g/mole

Lithium Dichromate



$$Li \quad 2 \times 7 = 14$$

$$Cr \quad 2 \times 52 = 104$$

$$O \quad 7 \times 16 = \underline{112}$$

230 g/mole

Barium Hydrogen Sulfate



$$Ba \quad 1 \times 137 = 137$$

$$H \quad 2 \times 1 = 2$$

$$S \quad 2 \times 32 = 64$$

$$O \quad 8 \times 16 = \underline{128}$$

331 g/mole

Gold I Thiosulfate



$$Au \quad 2 \times 197 = 394$$

$$S \quad 2 \times 32 = 64$$

$$O \quad 3 \times 16 = \underline{48}$$

506 g/mole