

## Naming simple ionic compounds

Rules: metal cation is always first, and is always the same name as the atom, the non-metal anion is always second.

If the anion is a mono-atomic anion, from the periodic table, the name ending changes to -ide. Oxygen becomes oxide, sulfur becomes sulfide, bromine becomes bromide (etc.).

No matter what the ratio of cations to anions is, we use the single metal name and the single anion name. No prefixes, such as di, tri, pent, or oct, are used in naming ionic compounds.

For each pair of metals and non metals, write the ions symbols with charges, write the neutral formula for the ionic compound that they form, then, name the compound.

Metal	Non-metal	Cation	Anion	Neutral compound formula	Compound name
Ba	S	Ba <sup>+2</sup>	S <sup>-2</sup>	BaS	Barium sulfide
Rb	N	Rb <sup>+1</sup>	N <sup>-3</sup>	Rb <sub>3</sub> N	Rubidium nitride
Li	O				
Al	F				
Sr	S				
Na	Br				
Ca	Cl				
Cs	P				
Be	O				
Mg	I				
K	O				
Na	P				