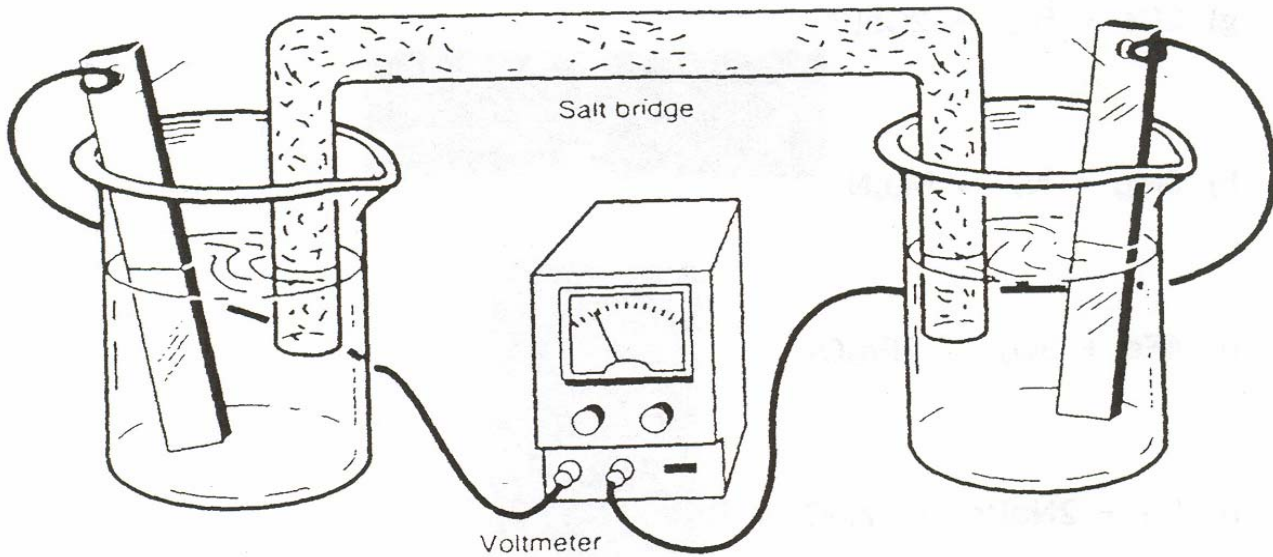


VOLTAIC CELL HANDOUT REDOX name: _____

Label the diagram below with the proper labels from the class demo. Answer the questions. Then, using the information provided on the back, label the second voltaic cell, and answer those questions.



1. Label the anode, cathode, flow of electrons, flow of both ions in salt bridge.
2. Describe the function of the salt bridge

3. Write the $\frac{1}{2}$ reactions:

$\frac{1}{2}$ Ox: _____ $\frac{1}{2}$ Red: _____

4. Write the NET REDOX reaction (no spectator ions, no electrons):

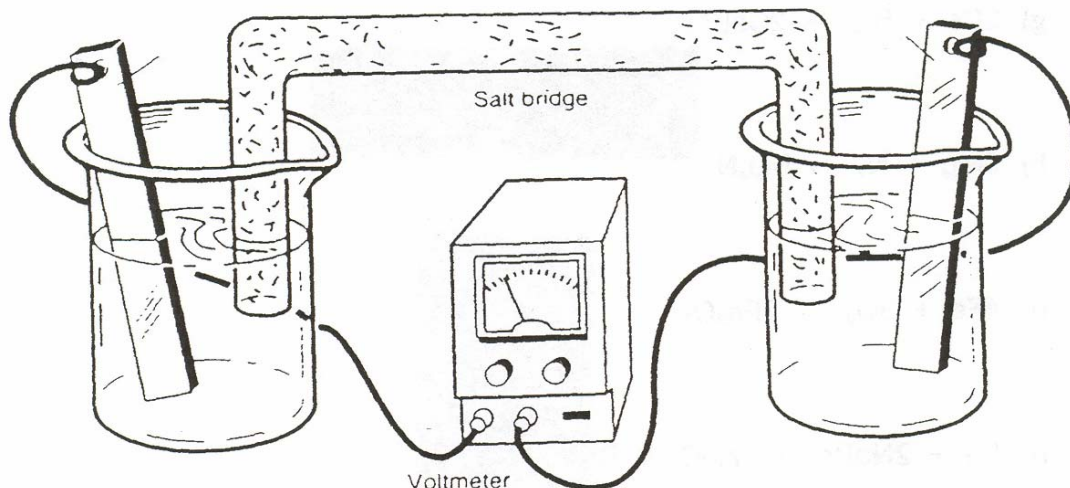
5. What is the reducing agent? _____ What is the oxidizing agent? _____

6. Which electrode loses mass? _____ where does it go? _____

7. A voltaic cell converts _____ into _____

If the metals are magnesium (right) and silver, the solutions are $Mg(NO_3)_2$ (at right) and $AgNO_3$ then label this diagram. The salt bridge is the same as above.

Note: On this page, the solution on the left side is the silver nitrate and the metal on the left is the silver.



8. Label the anode, cathode, flow of electrons, flow of both ions.

9. Describe the function of the salt bridge

10. Write the $\frac{1}{2}$ reactions:

$\frac{1}{2}$ Ox: _____ $\frac{1}{2}$ Red: _____

11. Write the NET REDOX reaction:

12. What is the reducing agent? _____ What is the oxidizing agent? _____

14. Which electrode gains mass? _____ how? _____

15. A voltaic cell converts _____ into _____