

Trends of the Periodic Table Atomic Sizes on the table and Atoms vs. Ions in size

Look at your periodic table before trying to answer this question on the back of this sheet. Then fill in the chart.

For the last column, write (for example) sulfur ion or sulfur atom

Why do atoms get bigger moving down a group and get smaller moving across the periods?

| name | symbol | atomic # | electron configuration | charge | which is bigger? |
|---------------|--------|----------|------------------------|--------|------------------|
| sulfur | | | | | |
| sulfur ion | | | | | |
| aluminum | | | | | |
| aluminum ion | | | | | |
| calcium | | | | | |
| calcium ion | | | | | |
| cesium | | | | | |
| cesium ion | | | | | |
| lithium | | | | | |
| lithium ion | | | | | |
| oxygen | | | | | |
| oxygen ion | | | | | |
| magnesium | | | | | |
| magnesium ion | | | | | |
| nitrogen | | | | | |
| nitrogen ion | | | | | |
| bromine | | | | | |
| bromine ion | | | | | |
| potassium | | | | | |
| potassium ion | | | | | |
| fluorine | | | | | |
| fluorine ion | | | | | |